

# Atomic Math



Name: \_\_\_\_\_

Date: \_\_\_\_\_ Class: \_\_\_\_\_ Seat #: \_\_\_\_\_

8
O
Oxygen
15.999

← \_\_\_\_\_

← \_\_\_\_\_

← \_\_\_\_\_

← \_\_\_\_\_

↓   ↓   ↓

needs to be changed to \_\_\_\_\_

← \_\_\_\_\_

← \_\_\_\_\_

**Atomic number** is the number of \_\_\_\_\_ .

**Atomic mass** is the AVERAGE number of \_\_\_\_\_ + \_\_\_\_\_ of all the isotopes of an atom .

**Mass number** is the number of \_\_\_\_\_ + \_\_\_\_\_ of a specific isotope of an atom .

8
O
_____
15.999

30
_____
Zinc
_____
65.39

3
Li
_____
6.941

Atomic # = \_\_\_\_\_

Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_

Atomic # = \_\_\_\_\_

Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

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# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_

Atomic # = \_\_\_\_\_

Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_

14
_____
Silicon
_____
28.086

5
B
_____
_____
10.81

35
_____
Bromine
_____
79.904

Atomic # = \_\_\_\_\_

Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_

Atomic # = \_\_\_\_\_

Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_

Atomic # = \_\_\_\_\_

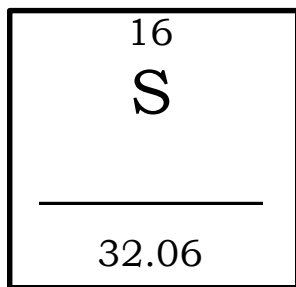
Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_



Atomic # = \_\_\_\_\_

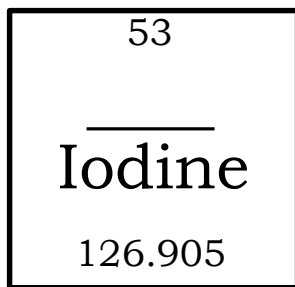
Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_



Atomic # = \_\_\_\_\_

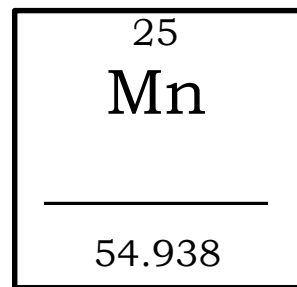
Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_



Atomic # = \_\_\_\_\_

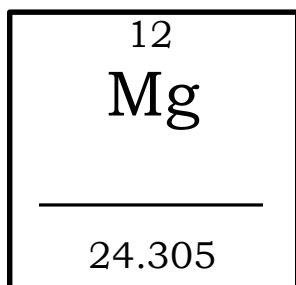
Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_



Atomic # = \_\_\_\_\_

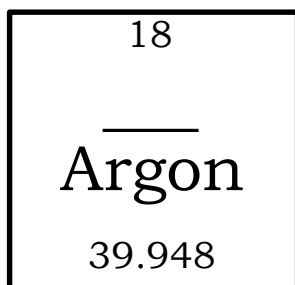
Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_



Atomic # = \_\_\_\_\_

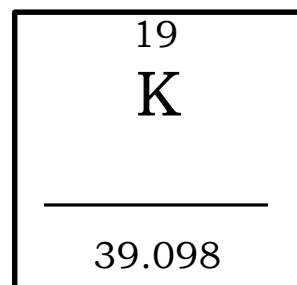
Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_



Atomic # = \_\_\_\_\_

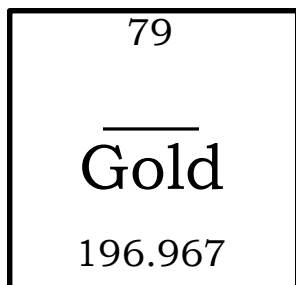
Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_



Atomic # = \_\_\_\_\_

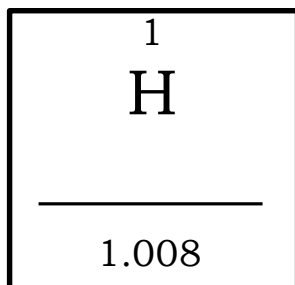
Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

# of Neutrons = \_\_\_\_\_

# of Electrons = \_\_\_\_\_



Atomic # = \_\_\_\_\_

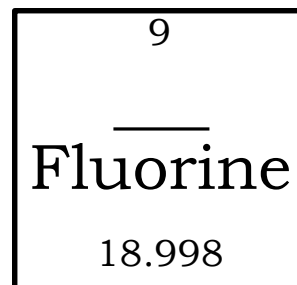
Atomic Mass = \_\_\_\_\_

Mass Number = \_\_\_\_\_

# of Protons = \_\_\_\_\_

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# of Electrons = \_\_\_\_\_



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